ADDITIONS AND CORRECTIONS

Notice to Readers.—For the convenience of those who wish to cut out the corrections and attach them to the margins of the articles corrected, they have been printed upon one side of the page only.

1941, Vol. 63

Louis F. Fieser and William H. Daudt. Syntheses in the Phenanthrene and Triphenylene Series.

Page 784. The names *dl*-α,β-dimethylsuccinic acid and *meso*-α,β-dimethylsuccinic acid on this page should be interchanged.—Louis F. Fieser.

1949, Vol. 71

C. G. Overberger, M. T. O'Shaughnessy and Harold Shalit. The Preparation of Some Aliphatic Azo Nitriles and their Decomposition in Solution.

Page 2666. In the **Summary**, col. 1, sect. 2, read "(R = ethyl..." instead of "(R = iso-butyl...."—C. G. Overberger.

1950, Vol. 72

David Edelson and Raymond M. Fuoss. A Contrast between Polyelectrolytes and Simple Electrolytes.

Page 309. In lines 9–11, the word "electrolyte" was accidentally duplicated. The phrase should read "the conductance of a simple electrolyte in the presence of a neutral chain polymer."—Raymond M. Fuoss.

C. G. Overberger, Evelyn J. Luhrs and P. K. Chien. The Preparation of *p*-Acyloxystyrenes.

Page 1200. A reference concerned with the application of the Perkin reaction to phenolic aldehydes was omitted: S. Ogawa, *Bull. Chem. Soc. Japan*, **2**, 20 (1927).—C. G. Overberger.

Carl M. Anderson, Leland G. Cole, and E. C. Gilbert. The Heats of Combustion of Some Isomeric Nitrostilbenes and their Isomerization Energies.

Page 1264. In col. 1, lines 9–10–11, the *trans*-4,4'-dinitrostilbenes melted at 303–304°, while the *cis*-isomer melted at 186.0–186.5°.—Leland G. Cole.

Noel F. Albertson. Piperidines and Azabicyclo Compounds. I. *Via* Michael Condensations.

Page 2596. In Table II footnote *j* should read, "m.p. > 275°" instead of 72–75°.—Noel F. Albertson.

D. W. Scott, H. L. Finke, W. N. Hubbard, J. P. McCullough, M. E. Gross, K. D. Williamson, Guy Waddington and H. M. Huffman. Spiropentane: Heat Capacity, Heats of Fusion and Vaporization, Vapor Pressure, Entropy and Thermodynamic Functions.

Page 4666. In col. 2, line 3 read

$$``B = +155 - 136.7e^{650/T} \text{ cc./mole} \qquad (3)"$$

and line 6, *et seq.*, read "At 283.16, 298.16 and 312.14° K. the values of *B* derived from the experimental data are −1217, −1042 and −943 cc./mole, respectively. These agree well with the values −1201, −1055 and −941 calculated from equation 3."—Guy Waddington.

Nelson K. Richtmyer, Laura C. Stewart and C. S. Hudson. L-Fuco-4-ketose, a New Sugar Produced by the Action of *Acetobacter suboxydans* on L-Fucitol.

Page 4936. In col. 2, lines 34 and 54, for "2,4:3,5-" read "1,3:2,4-."—Nelson K. Richtmyer.

1951, Vol. 73

Stanley J. Cristol, John R. Douglass and John S. Meek. Reactions of Organometallic Compounds. III. The Reactions of Phenyllithium with Some Epoxides.

Page 817. In footnote (25) read "Danilov" for "Danislov" and read "*Chem. Zentr.*, **94**, III, 761 (1923)" for "*Chem. Zentr.*, 94, II, 761 (1923)."—Stanley J. Cristol.

John Weijlard, Karl Pfister, 3rd, Edward F. Swanezy, C. A. Robinson and Max Tishler. Preparation of the Stereoisomeric α,β-Diphenyl-β-hydroxyethylamines.

Page 1217. In col. 9, lines 4 and 5, and col. 2, line 17, for "D-glutamic" read "L-(+)-glutamic." In col. 2, lines 15 and 16, lines 37 and 38, for "D-*erythro*-α,β-Diphenyl-β-hydroxyethylamine-D-glutamate" read "D-*erythro*-α,β-Diphenyl-β-hydroxyethylamine-L-glutamate." In col. 2, lines 36 and 37, lines 48 and 49, for "L-*erythro*-α,β-Diphenyl-β-hydroxyethylamine-D-glutamate" read "L-*erythro*-α,β-Diphenyl-β-hydroxyethylamine-L-glutamate".—Karl Pfister, 3rd.

Norman H. Cromwell, David B. Capps and S. Edward Palmer. Endocyclic α,β-Unsaturated Ketones. II. Reactions of 8-Bromoperinaphthenone-7 with Amines.

Page 1228. In Table I. Interchange the first listed values under V and VI. First recorded value for compound VI to now read, 3357 cm⁻¹, 30% Approx. % abs.; for compound V to now read, 3045 cm⁻¹, 2–3% Approx. % abs.—Norman H. Cromwell.

Herbert E. Ungnade and Ishmael Ortega. The Absorption Spectra of Acrylic Acids and Esters.

Page 1565. In Col. 1, variations of extinctions, ascribed to failure of Beer's law, are due to instrumental deviations only. The authors are indebted to Dr. L. J. Saidel for pointing out that a larger density range was used than is advisable for optimum accuracy. Since concentrations were given, it is fairly easy to recognize and discard inaccurate values.

In view of Dr. J. Cason's report of acid spectra in which constant slit width is mandatory (*J. Org. Chem.*, **16**, 1180 (1951)) and recent reports of apparent Beer's law failure near the transmission limit of ethanol (*Science*, **114**, 576 (1951)), it appeared advisable to recheck one of our acids. The results with β,β-dimethylacrylic acid in the following table show that the extinction values at the maximum vary slightly with slit width and density but are little affected by cell length and concentration.

Concentration moles/l. × 10^5	Density range	Cell length, cm.	Slit, mm.	$\lambda_{\text{max.}}$,[d] mμ	log ε
5.75	0.1–0.7	1.00	*a*	216	4.10
6.26	0.4–0.8	1.00	1.0	216	4.10[b]
13.3	1.3–1.7	1.00	1.0	216	4.11
31.3	0.2–0.5	0.11	1.0	216	4.13
66.8	0.7–1.0	0.11	1.0	216	4.13
71.0	0.7–1.0	0.11	[a]	216	4.10[c]

[a] Narrowest slit obtainable. [b] For slit variation of 2.0–0.6 mm. (at 216 mμ), d., 0.800–0.802. [c] Log ϵ_{240} 3.21–3.27 at $5.7 \times 10^{-5} - 7.1 \times 10^{-4}$ moles/l. (d., 0.09–0.54). [d] All values below 220 mμ were determined with the 10,000 megohm resistor which increases the sensitivity of the instrument five-fold.

Page 1566. Add a footnote: "diluted 1:10 at the maximum." to the highest concentration of compound IV in Table II.—H. E. Ungnade.

Leland J. Kitchen. Nopadiene.

Page 2369. In col. 1, line 16 from the end, for "$\alpha^{25}\text{D}$ −30.16°" read "$\alpha^{25}\text{D}$ +59.84°."—L. J. Kitchen.